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CHEMICAL METALS AND THEIR WORKING

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ANNOTATION
This article basically shows the process of working with chemicals and solutions.
Key words: their composition is changeable, mechanical mixture

INTRODUCTION
The homogeny system, consisting of two or more components and their interaction products, is called a solution. Solutions are important in the life of living organisms. For example, blood, lymph, and saliva fluids are solutions. Solutions occupy an intermediate position between chemical compounds and simple mechanical compounds. Solutions are homogeneous, that is, homogen, similar to chemical compounds. The melting process occurs with the absorption or release of heat, similar to a chemical reaction. Solutions differ from chemical compounds in that their composition is changeable, showing both the properties of the solvent and the dissolved substance. The dissolved substance can be separated from the Solvent by a physical way. With these properties, the solution is similar to a mechanical mixture.

In the process of preparation of the solution, the aggregate state is a variable component solvent. In solutions, one component in the other, so as to form a homogeneous environment in the state of molecules, ions, or atoms spread. Solutions three in gas, liquid and solid state.

METHODS OF EXPRESSING THE CONCENTRATION OF SOLUTIONS
The amount of substance dissolved in a unit of volume or mass of a solution or solvent is called concentration. Such a solution is a concentrated, slightly diluted solution, if there is a lot of dissolved substance in the solution. There are several ways of expressing concentration. 1. The ratio of the mass of the dissolved substance to the total mass of the solution indicates the mass fraction of the dissolved substance:

Metals (Yun. metalleuo — I dig, I dig from the ground) - simple substances that have such distinctive properties as high electrical conductivity, hot conductivity, electrical conductivity, good return of electric magnetic waves, plasticity, under normal conditions. M. in the solid state, the crystal is in the structure. In Steam, it is an atom. oxides in combination with water often turn into hydroxides (bases). M. the electronic structure has its own characteristics, which are mentioned above in the true-file. M. atoms easily give external (valence) electrons. The crystal lattice, not all electrons will be combined with their atoms. Some of them move.
Chemical properties. D. I. 109 of the 87 chemical elements in Mendeleev's process system M., 22 of them are metallic. All M.ni it forms" ordinary metals"," intermediate metals"," lantanoids and actinoids". Process in the system, metals in the main groups are called ordinary metals (s - and r-elements),

metals in the additional Group — Intermediate metals or (D - and f - elements). Simple substances are conditionally divided into two groups as metals and. Mas, Ge and SB do not have a single opinion about which category. But since Germany has semiconductor properties, it is not, although is a semi-metal according to its physical properties, M. it is more correct to count as. Tin has metal (Z-Sn) and semi-conductor (a-Sn) modifications. Germany, Silicon, phosphorus and some under high pressure M. as determined by the presence of conductive modifications. In addition, all substances under high pressure can also exhibit metallic properties. Therefore, it is necessary to use this or that element M.ga or when determining whether it is resistant to, it is necessary to take into account not only its physical properties, but also its chemical properties. M. it enters chemical reactions as an electron donor, forming positively charged ions in compounds or solutions. M.on electromagnetism is lower than the electromagnetism. Many M. it reacts actively with hydrogen, halogen. Alkaline and alkaline earth metals at normal temperatures with water, m as zinc and iron. esa reacts with water vapor at high temperatures. Series M with nitrogen., mas, lithium reacts at room temperature, magnesium, zirconium, gaily, titanium when heated. The metal itself compresses the original metal from that metal solution.

Density is smaller than 5 M. light, older than 5 years-heavy M. it is called. Iron and its alloys are black M., the rest is colored M. refers to as. The Original M. he does not look at it. Rare M. in the sentence vanadium, molybdenum, beryllium, indium, zirconium, lanthanum, niobium, tantalum, renium, German, galley, thallium and others are included. "Rare M.the phrase " quot; is a conditional phrase, depends on how much the methods of separating pure metal are improved; titanium, which was once considered "rare", now does not refer to the sentence" rare " (M.on the chemical and physical properties of metals, see articles on metal elements). M.in some compounds (even in alloys), metal bonding (bonding between particles that form the metal) is preserved. In the case of free and chemical compounds in nature. The Original M. (gold, platinum, silver), sometimes copper, tin and Mercury are found in pure form.

Ores M.ni the work of obtaining in pure form is carried out in various branches of metallurgy (drilling, hydrometallurgy and electric metallurgy)as a result of the processes of return, thermal decomposition, exchange in the technique. Very pure M. to obtain substances, the method of driving in a vacuum is also used. In subsequent years, the method of liquefaction along the zones is often used. On the basis of this method, niobium, tantalum, tungsten and other M. the yacht is cleaned of substances. M. in pure form, it is rarely used. Most often, it is used in the case of an alloy. Mas, cast iron, steel, brass, bronze, Constantine, milkier, Nixon and other atmospheric conditions M. edible (corrosion). It is of paramount importance to store metal objects from decay. The finding of a method of preparation of special stainless steels will help to solve this issue. M. it is used in marriage, construction, Cosmonautics, Shipbuilding, Mechanical Engineering, aircraft building and many other fields (neither q nor Q. Working metals).

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